

# ANSWERS

## Chapter 8

- 8.1** 1.8
- 8.2** (a) From the given graph for a stress of  $150 \times 10^6 \text{ N m}^{-2}$  the strain is 0.002  
(b) Approximate yield strength of the material is  $3 \times 10^8 \text{ N m}^{-2}$
- 8.3** (a) Material A  
(b) Strength of a material is determined by the amount of stress required to cause fracture: material A is stronger than material B.
- 8.4** (a) False (b) True
- 8.5**  $1.5 \times 10^{-4} \text{ m}$  (steel);  $1.3 \times 10^{-4} \text{ m}$  (brass)
- 8.6** Deflection =  $4 \times 10^{-6} \text{ m}$
- 8.7**  $2.8 \times 10^{-6}$
- 8.8** 0.127
- 8.9**  $7.07 \times 10^4 \text{ N}$
- 8.10**  $D_{\text{copper}}/D_{\text{iron}} = 1.25$
- 8.11**  $1.539 \times 10^{-4} \text{ m}$
- 8.12**  $2.026 \times 10^9 \text{ Pa}$
- 8.13**  $1.034 \times 10^3 \text{ kg/m}^3$
- 8.14** 0.0027
- 8.15**  $0.058 \text{ cm}^3$
- 8.16**  $2.2 \times 10^6 \text{ N/m}^2$

## Chapter 9

- 9.3** (a) decreases (b)  $\eta$  of gases increases,  $\eta$  of liquid decreases with temperature (c) shear strain, rate of shear strain (d) conservation of mass, Bernoulli's equation (e) greater.
- 9.5**  $6.2 \times 10^6$  Pa
- 9.6** 10.5 m
- 19.7** Pressure at that depth in the sea is about  $3 \times 10^7$  Pa. The structure is suitable since it can withstand far greater pressure or stress.
- 9.8**  $6.92 \times 10^5$  Pa
- 9.9** 0.800
- 9.10** Mercury will rise in the arm containing spirit; the difference in levels of mercury will be 0.221 cm.
- 9.11** No, Bernoulli's principle applies to streamline flow only.
- 9.12** No, unless the atmospheric pressures at the two points where Bernoulli's equation is applied are significantly different.
- 9.13**  $9.8 \times 10^2$  Pa (The Reynolds number is about 0.3 so the flow is laminar).
- 9.14**  $1.5 \times 10^3$  N
- 9.15** Fig (a) is incorrect [Reason: at a constriction (i.e. where the area of cross-section of the tube is smaller), flow speed is larger due to mass conservation. Consequently pressure there is smaller according to Bernoulli's equation. We assume the fluid to be incompressible].
- 9.16**  $0.64 \text{ m s}^{-1}$
- 9.17**  $2.5 \times 10^{-2} \text{ N m}^{-1}$
- 9.18**  $4.5 \times 10^{-2}$  N for (b) and (c), the same as in (a).
- 9.19** Excess pressure = 310 Pa, total pressure =  $1.0131 \times 10^5$  Pa. However, since data are correct to three significant figures, we should write total pressure inside the drop as  $1.01 \times 10^5$  Pa.
- 9.20** Excess pressure inside the soap bubble = 20.0 Pa; excess pressure inside the air bubble in soap solution = 10.0 Pa. Outside pressure for air bubble =  $1.01 \times 10^5 + 0.4 \times 10^3 \times 9.8 \times 1.2 = 1.06 \times 10^5$  Pa. The excess pressure is so small that up to three significant figures, total pressure inside the air bubble is  $1.06 \times 10^5$  Pa.

## Chapter 10

- 10.1** Neon:  $-248.58 \text{ }^\circ\text{C} = -415.44 \text{ }^\circ\text{F}$ ;  
CO<sub>2</sub>:  $-56.60 \text{ }^\circ\text{C} = -69.88 \text{ }^\circ\text{F}$

$$\text{(use } t_{\text{F}} = \frac{9}{5}t_{\text{C}} + 32)$$

- 10.2**  $T_A = (4/7) T_B$
- 10.3** 384.8 K
- 10.4** (a) Triple-point has a *unique* temperature; fusion point and boiling point temperatures depend on pressure; (b) The other fixed point is the absolute zero itself; (c) Triple-point is  $0.01^\circ\text{C}$ , not  $0^\circ\text{C}$ ; (d) 491.69.
- 10.5** (a)  $T_A = 392.69\text{ K}$ ,  $T_B = 391.98\text{ K}$ ; (b) The discrepancy arises because the gases are not perfectly ideal. To reduce the discrepancy, readings should be taken for lower and lower pressures and the plot between temperature measured versus absolute pressure of the gas at triple point should be extrapolated to obtain temperature in the limit pressure tends to zero, when the gases approach ideal gas behaviour.
- 10.6** Actual length of the rod at  $45.0^\circ\text{C} = (63.0 + 0.0136)\text{ cm} = 63.0136\text{ cm}$ . (However, we should say that change in length up to three significant figures is  $0.0136\text{ cm}$ , but the total length is  $63.0\text{ cm}$ , up to three significant places. Length of the same rod at  $27.0^\circ\text{C} = 63.0\text{ cm}$ .
- 10.7** When the shaft is cooled to temperature  $-69^\circ\text{C}$  the wheel can slip on the shaft.
- 10.8** The diameter increases by an amount  $= 1.44 \times 10^{-2}\text{ cm}$ .
- 10.9**  $3.8 \times 10^2\text{ N}$
- 10.10** Since the ends of the combined rod are not clamped, each rod expands freely.  
 $\Delta l_{\text{brass}} = 0.21\text{ cm}$ ,  $\Delta l_{\text{steel}} = 0.126\text{ cm} = 0.13\text{ cm}$   
Total change in length  $= 0.34\text{ cm}$ . No 'thermal stress' is developed at the junction since the rods freely expand.
- 10.11**  $0.0147 = 1.5 \times 10^{-2}$
- 10.12**  $103^\circ\text{C}$
- 10.13**  $1.5\text{ kg}$
- 10.14**  $0.43\text{ J g}^{-1}\text{ K}^{-1}$ ; smaller
- 10.15** The gases are diatomic, and have other degrees of freedom (i.e. have other modes of motion) possible besides the translational degrees of freedom. To raise the temperature of the gas by a certain amount, heat is to be supplied to increase the average energy of all the modes. Consequently, molar specific heat of diatomic gases is more than that of monatomic gases. It can be shown that if only rotational modes of motion are considered, the molar specific heat of diatomic gases is nearly  $(5/2) R$  which agrees with the observations for all the gases listed in the table, except chlorine. The higher value of molar specific heat of chlorine indicates that besides rotational modes, vibrational modes are also present in chlorine at room temperature.
- 10.16**  $4.3\text{ g/min}$
- 10.17**  $3.7\text{ kg}$
- 10.18**  $238^\circ\text{C}$
- 10.20**  $9\text{ min}$

## Chapter 11

- 11.1** 16 g per min
- 11.2** 934 J
- 11.4** 2.64
- 11.5** 16.9 J
- 11.6** (a) 0.5 atm (b) zero (c) zero (assuming the gas to be ideal) (d) No, since the process (called free expansion) is rapid and cannot be controlled. The intermediate states are non-equilibrium states and do not satisfy the gas equation. In due course, the gas does return to an equilibrium state.
- 11.7** 25 W
- 11.8** 450 J

## Chapter 12

- 12.1**  $4 \times 10^{-4}$
- 12.3** (a) The dotted plot corresponds to 'ideal' gas behaviour; (b)  $T_1 > T_2$ ; (c)  $0.26 \text{ J K}^{-1}$ ; (d) No,  $6.3 \times 10^{-5} \text{ kg}$  of  $\text{H}_2$  would yield the same value
- 12.4** 0.14 kg
- 12.5**  $5.3 \times 10^{-6} \text{ m}^3$
- 12.6**  $6.10 \times 10^{26}$
- 12.7** (a)  $6.2 \times 10^{-21} \text{ J}$  (b)  $1.24 \times 10^{-19} \text{ J}$  (c)  $2.1 \times 10^{-16} \text{ J}$
- 12.8** Yes, according to Avogadro's law. No,  $v_{\text{rms}}$  is largest for the lightest of the three gases; neon.
- 12.9**  $2.52 \times 10^3 \text{ K}$
- 12.10** Use the formula for mean free path :

$$\bar{l} = \frac{1}{\sqrt{2} \pi n d^2}$$

where  $d$  is the diameter of a molecule. For the given pressure and temperature  $N/V = 5.10 \times 10^{25} \text{ m}^{-3}$  and  $d = 1.0 \times 10^{-7} \text{ m}$ .  $v_{\text{rms}} = 5.1 \times 10^2 \text{ m s}^{-1}$ .

collisional frequency =  $\frac{v_{\text{rms}}}{\lambda} = 5.1 \times 10^9 \text{ s}^{-1}$ . Time taken for the collision =  $d / v_{\text{rms}} = 4 \times 10^{-13} \text{ s}$ .

Time taken between successive collisions =  $1 / v_{\text{rms}} = 2 \times 10^{-10} \text{ s}$ . Thus the time taken between successive collisions is 500 times the time taken for a collision. Thus a molecule in a gas moves essentially free for most of the time.

## Chapter 13

**13.1** (b), (c)

**13.2** (b) and (c): SHM; (a) and (d) represent periodic but not SHM [A polyatomic molecule has a number of natural frequencies; so in general, its vibration is a superposition of SHM's of a number of different frequencies. This superposition is periodic but not SHM].

**13.3** (b) and (d) are periodic, each with a period of 2 s; (a) and (c) are not periodic. [Note in (c), repetition of merely one position is not enough for motion to be periodic; the entire motion during one period must be repeated successively].

**13.4** (a) Simple harmonic,  $T = (2\pi/\omega)$ ; (b) periodic,  $T = (2\pi/\omega)$  but not simple harmonic; (c) simple harmonic,  $T = (\pi/\omega)$ ; (d) periodic,  $T = (2\pi/\omega)$  but not simple harmonic; (e) non-periodic; (f) non-periodic (physically not acceptable as the function  $\rightarrow \infty$  as  $t \rightarrow \infty$ ).

**13.5** (a) 0, +, +; (b) 0, -, -; (c) -, 0, 0; (d) -, -, -; (e) +, +, +; (f) -, -, -.

**13.6** (c) represents a simple harmonic motion.

**13.7**  $A = \sqrt{2} \text{ cm}$ ,  $\phi = 7\pi/4$ ;  $B = \sqrt{2} \text{ cm}$ ,  $a = \pi/4$ .

**13.8** 219 N

**13.9** Frequency  $3.2 \text{ s}^{-1}$ ; maximum acceleration of the mass  $8.0 \text{ m s}^{-2}$ ; maximum speed of the mass  $0.4 \text{ m s}^{-1}$ .

**13.10** (a)  $x = 2 \sin 20t$   
 (b)  $x = 2 \cos 20t$   
 (c)  $x = -2 \cos 20t$

where  $x$  is in cm. These functions differ neither in amplitude nor frequency. They differ in initial phase.

**13.11** (a)  $x = -3 \sin \pi t$  where  $x$  is in cm.

(b)  $x = -2 \cos \frac{\pi}{2} t$  where  $x$  is in cm.

**13.13** (a)  $F/k$  for both (a) and (b).

(b)  $T = 2\pi\sqrt{\frac{m}{k}}$  for (a) and  $2\pi\sqrt{\frac{m}{2k}}$  for (b)

**13.14** 100 m/min

**13.15** 8.4 s

**13.16**  $T = 2\pi \sqrt{\frac{l}{\sqrt{g^2 + v^4/R^2}}}$ . Hint: Effective acceleration due to gravity will get reduced due to radial acceleration  $v^2/R$  acting in the horizontal plane.

**13.17** In equilibrium, weight of the cork equals the up thrust. When the cork is depressed by an amount  $x$ , the net upward force is  $Ax\rho_1g$ . Thus the force constant  $k = A\rho_1g$ .

Using  $m = Ah\rho$ , and  $T = 2\pi \sqrt{\frac{m}{k}}$  one gets the given expression.

**13.18** When both the ends are open to the atmosphere, and the difference in levels of the liquid in the two arms is  $h$ , the net force on the liquid column is  $Ah\rho g$  where  $A$  is the area of cross-section of the tube and  $\rho$  is the density of the liquid. Since restoring force is proportional to  $h$ , motion is simple harmonic.

## Chapter 14

**14.1** 0.5 s

**14.2** 8.7 s

**14.3**  $2.06 \times 10^4 \text{ N}$

**14.4** Assume ideal gas law:  $P = \frac{\rho RT}{M}$ , where  $\rho$  is the density,  $M$  is the molecular mass, and

$T$  is the temperature of the gas. This gives  $v = \sqrt{\frac{\gamma RT}{M}}$ . This shows that  $v$  is:

- (a) Independent of pressure.
- (b) Increases as  $\sqrt{T}$ .
- (c) The molecular mass of water (18) is less than that of  $\text{N}_2$  (28) and  $\text{O}_2$  (32).

Therefore as humidity increases, the effective molecular mass of air decreases and hence  $v$  increases.

**14.5** The converse is not true. An obvious requirement for an acceptable function for a travelling wave is that it should be finite everywhere and at all times. Only function (c) satisfies this condition, the remaining functions cannot possibly represent a travelling wave.

**14.6** (a)  $3.4 \times 10^{-4} \text{ m}$       (b)  $1.49 \times 10^{-3} \text{ m}$

- 14.7**  $4.1 \times 10^{-4}$  m
- 14.8** (a) A travelling wave. It travels from right to left with a speed of  $20 \text{ ms}^{-1}$ .  
(b) 3.0 cm, 5.7 Hz  
(c)  $\pi/4$   
(d) 3.5 m
- 14.9** All the graphs are sinusoidal. They have same amplitude and frequency, but different initial phases.
- 14.10** (a)  $6.4 \pi$  rad  
(b)  $0.8 \pi$  rad  
(c)  $\pi$  rad  
(d)  $(\pi/2)$  rad
- 14.11** (a) Stationary wave  
(b)  $l = 3$  m,  $n = 60$  Hz, and  $v = 180 \text{ m s}^{-1}$  for each wave  
(c) 648 N
- 14.12** (a) All the points except the nodes on the string have the same frequency and phase, but not the same amplitude.  
(b) 0.042 m
- 14.13** (a) Stationary wave.  
(b) Unacceptable function for any wave.  
(c) Travelling harmonic wave.  
(d) Superposition of two stationary waves.
- 14.14** (a)  $79 \text{ m s}^{-1}$   
(b) 248 N
- 14.15**  $347 \text{ m s}^{-1}$
- Hint :  $v_n = \frac{(2n-1)v}{4l}$  ;  $n = 1, 2, 3, \dots$  for a pipe with one end closed
- 14.16**  $5.06 \text{ km s}^{-1}$
- 14.17** First harmonic (fundamental); No.
- 14.18** 318 Hz

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## NOTES

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